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Which one of the following reactions is NOT balanced?

$2 \text{CO} + \text{O}_2 \rightarrow 2 \text{CO}_2$

$2 \text{SO}_2 + \text{O}_2 \rightarrow 2 \text{SO}_3$

$\text{SF}_4 + 3 \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_3 + 4 \text{HF}$

$2 \text{KNO}_3 \rightarrow 2 \text{K}_2\text{O} + \text{N}_2$

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ANSWER: Correct

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### Chapter 9 Reading

#### Question 2 Part A Iron

reacts with oxygen in the air to produce rust ( $\text{Fe}_2\text{O}_3$ ) according to the reaction shown

below.  $4 \text{Fe} + 3 \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3$

How many moles of rust can be formed if 1.8 mols of iron are present? Hint 1. Begin with your given number of moles.

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ANSWER Answer the following questions in the space provided. 1. List four metals that will not replace hydrogen in an acid. Choose from Cu, Ag, Au, Pt, Sb, Bi, and Hg.

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Chapter 2: Rank from  
most to fewest valence



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electrons of the following elements: Kr, Ga, Br, As, K. - Answer: Kr (8), Br (7), As (5), Ga (3), K (1). Valence electrons are equal to the atom's main group number. Link: <https://www.coursehero.com/file/20333255/Mastering-Chemistry-Chapter-2/>.

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