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0.45 mol O₂. c. 2

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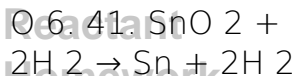
13.5 g O₂. d. 4

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45.654 g C₆H₁₂

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O. Limiting reactant: 5.6 mol H

2. 3.3×10^2 g Sn,

11 g H₂, $1.0 \times$

10^2 g H₂O, 90.

g SnO₂. Challenge

Problems. 43. $6\text{CO}_2 + 6\text{H}_2\text{O} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2$.

57.5 g C₆H₁₂O₆

6. 45. 81% CO₂.

57% H₂O. 90%

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$\text{NaC}_2\text{H}_3\text{O}_2$.

Limiting reactant:

0.99 g NaHCO_3

0.52 g CO_2 , 0.21 g

H_2O , 0.97 g $\text{NaC}_2\text{H}_3\text{O}_2$.

$\text{Mg}(\text{OH})_2$.

$\text{Mg}(\text{OH})_2(\text{s}) + 2\text{HCl}$

$(\text{aq}) \rightarrow \text{MgCl}_2 \dots$

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sodium hydroxide

was the limiting

reactant and 48.64

grams of sodium

phosphate is

formed. To

determine the

amount of excess

reactant remaining,

the amount used is

needed. grams of

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reactant used =
(grams of product
formed) x (1 mol of
product/molar
mass of product) x
(mole ratio of
reactant/product) x
(molar mass of
reactant)

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